**Ionic and Covalent Review Student Worksheet #1**

*Please write the formula or name corresponding to the given name or formula for the following ionic compounds. You can use the “subscript”* *button in the “Home” bar to make small numbers.*

*Hints:*

1. *For every question ask yourself, “Is this covalent or ionic?”.*
2. *If the question is covalent, use prefixes and DO NOT BALANCE charges. If you need help, use the flow chart for covalent compounds found on the back of the chemistry reference package.*
3. *If the question is ionic, you need to draw a t-chart to balance the charges. DO NOT use prefixes. If you need help, use the flow chart for ionic compounds found on the back of the chemistry reference package.*
4. *If an element ends in something other than ‘ide’ it might be polyatomic. Treat these as ionic compounds. We have created these in* ***bold*** *below.*
5. *We recommend you do this on paper, take a picture of your work, and submit it in TEAMS just below the word document where it says: “My Work + work.”*

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| Sodium bromide | NaCl |
| Aluminum fluroride | Cs3P |
| Tetraselinium monooxide | N2Br4 |
| Silver nitride | CuI |
| Gold (III) iodide | (NH4)2SO3 |
| Telluride dinitride | C5F |

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| --- | --- |
| Aluminium chrom**ate** | Mn(CN)2 |
| Trisulfur nonaoxide | KCH3C00 |
| Ammonium nitr**ite** | Se5Cl2 |
| Manganese (II) phosph**ite** | Pb(OH)4 |
| Calcium sulphide | CuI2 |
| Lead (IV) sulph**ate** | Fe3(PO3)2 |